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Total \u0026 Net Ionic  
Equations Acid Base  
Neutralization Reactions  
\u0026 Net Ionic Equations -  
*Chemistry Chemistry 1a Net  
Ionic Answer*

ionic:  $\text{H}^+ (\text{aq}) + \text{Cl}^- (\text{aq}) + \text{K}^+ (\text{aq}) + \text{OH}^- (\text{aq}) \rightarrow \text{K}^+ (\text{aq}) +$

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$\text{Cl}^-(\text{aq}) + \text{H}_2\text{O}(\text{l})$  net  
ionic:  $\text{H}^+(\text{aq}) + \text{OH}^-(\text{aq}) \rightarrow \text{H}_2\text{O}(\text{l})$

3.  $\text{H}_2\text{SO}_4(\text{aq}) + \text{Mg}(\text{OH})_2(\text{s}) \rightarrow \text{MgSO}_4(\text{aq}) + 2\text{H}_2\text{O}(\text{l})$   
ionic:  $2\text{H}^+(\text{aq}) + \text{SO}_4^{2-}(\text{aq}) + \text{Mg}^{2+}(\text{aq}) + 2\text{OH}^-(\text{aq}) \rightarrow \text{Mg}^{2+}(\text{aq}) + \text{SO}_4^{2-}(\text{aq}) + 2\text{H}_2\text{O}(\text{l})$  net

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ionic:  $\text{H}^+ (\text{aq}) + \text{OH}^- (\text{aq}) \rightarrow \text{H}_2\text{O} (\text{l})$

4.  $\text{NH}_4\text{OH} (\text{aq}) + \text{HCl} (\text{aq}) \rightarrow \text{NH}_4\text{Cl} (\text{aq}) + \text{H}_2\text{O} (\text{l})$

ionic:  $\text{NH}_4^+ (\text{aq}) + \text{OH}^- (\text{aq}) + \text{H}^+ (\text{aq}) + \text{Cl}^- (\text{aq}) \rightarrow \text{NH}_4^+ (\text{aq}) + \text{H}_2\text{O} (\text{l}) + \text{Cl}^- (\text{aq})$

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*Equations Answers*

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*Task#1A . . .*

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different chloride solutions  
are mixed with each other:

Solution 1 is an aqueous  
solution of 4.85 g aluminum  
chloride in 150. ml water

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and solution 2 has a volume of 175 ml and contains 7.75 g of zinc chloride. You mix the two solutions together in a large beaker. Solved:  
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Experimental Errors,  
Significant Figures, and Net  
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Name \_\_\_\_\_ Date \_\_\_\_\_

5. Iron III nitrate and potassium sulfide I

6. strontium acetate and phosphoric acid

- 7. aluminum sulfate and ammonium sulfide

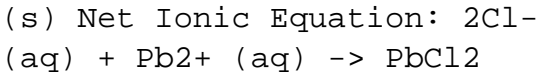
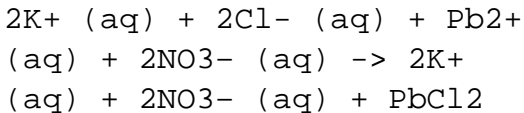
8. barium hydroxide and zinc(II)

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(s) Directions: Write balanced molecular, ionic, and net ionic equations for each of the following reactions. Assume all reactions occur in aqueous solution.

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*Net Ionic Equation Worksheet  
Answers*

Chemistry 1a Net Ionic  
Answer Answer to 1a.

Identify whether each  
species functions as a  
Brønsted-Lowry acid or a  
Brønsted-Lowry base in this



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## Net Ionic Answer Key

net ionic equati... Solved:  
1a. Identify Whether Each  
Species Functions As A B ...  
 $2\text{K}^+ (\text{aq}) + 2\text{Cl}^- (\text{aq}) + \text{Pb}^{2+}$   
 $(\text{aq}) + 2\text{NO}_3^- (\text{aq}) \rightarrow 2\text{K}^+$   
 $(\text{aq}) + 2\text{NO}_3^- (\text{aq}) + \text{PbCl}_2$   
(s) Net Ionic

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ionic equation for the  
reaction that occurs when



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barium sulfite (s) and  
excess nitric acid (aq) are  
combined. Note: Sulfites  
follow the same solubility  
trends as sulfates.

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