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Solving the chemical equilibria problem boils down to finding the extents of reaction and temperature that equate the equilibrium coefficient calculated both ways. Consider the reaction $A + B \rightleftharpoons C + D$ (12) with initial mole fractions of each species and moles of each species given by

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Chemical equilibrium, a condition in the course of a reversible chemical reaction in which no net change in the amounts of reactants and products occurs. A reversible chemical reaction is one in which the products, as soon as they are formed, react to produce the original reactants. At equilibrium, the two opposing reactions go on at equal rates, or velocities, hence there is no net change in the amounts of substances involved.

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Given the equation, $N_2(g) + 3H_2(aq) \rightleftharpoons 2NH_3(g)$, Find Q and determine which direction the reaction will shift in order to reach the state of chemical equilibrium. Solution: Given, $[N_2] = 0.04M$, $[H_2] = 0.09M$, and $K = 0.040$

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Third, we have chemical equilibrium, equating the chemical species of each comonent in the two phases, $V_L \mu^? = \mu^?$ for $? = 1$ to N_c , (3)

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Chemical equilibrium is the state of a reversible reaction where the rate of the forward reaction equals the rate of the reverse reaction. While a reaction is in equilibrium the concentration of the reactants and products are constant.

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